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Total



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1. Complete the table.

Name of compound	Formula of compound	Formula of cation	Formula of anion
magnesium chloride	MgCl_2	Mg^{2+}	Cl ⁻
lead(II) nitrate		Pb ²⁺	
aluminium sulphate		Al ³⁺	SO ₄ ²⁻
manganese(III) oxide	Mn_2O_3		O ²⁻
calcium phosphate	Ca ₃ (PO ₄) ₂		

Q1

(Total 6 marks)

JIV	e the name of each of the following substances.	
(a)	The gas that occupies about 79% of the air by volume.	
		(1)
(b)	A dark grey solid that sublimes to form a violet vapour.	
		(1)
(c)	A non-metallic element that conducts electricity.	
,		(1)
(d)	A yellow element that exists as allotropes.	
(4)	J • ·· • • • • • • •	(1)
(e)	A blue solid that decomposes on heating to evolve a brown gas.	
(0)	71 blue solid that decomposes on heating to evolve a blown gas.	
		(1)
(f)	The gas that is given off when sodium is added to ethanol.	
		(1)
		al 6 marks)

State two methods which are used to prevent the rusting of iron. (2) The diagram below shows an iron nail in contact with a strip of magnesium ribbon. Explain whether the rate of rusting would be increased, decreased or unchanged by the presence of the magnesium. Water
The diagram below shows an iron nail in contact with a strip of magnesium ribbon. Explain whether the rate of rusting would be increased, decreased or unchanged by the presence of the magnesium. Water
The diagram below shows an iron nail in contact with a strip of magnesium ribbon. Explain whether the rate of rusting would be increased, decreased or unchanged by the presence of the magnesium. Water
Explain whether the rate of rusting would be increased, decreased or unchanged by the presence of the magnesium. Water Nail
Nail
Nail
Nail
Nail
Magnesium ribbon
The rate of rusting is because
(3)
(Total 6 marks)

Leave blank **4.** Complete the table.

Particle	Number of protons	Number of neutrons	Number of electrons
$_{30}^{65}$ Zn	30		
$^{39}_{19}{ m K}^+$		20	
	16	16	18

	19K+		20	
		16	16	18
		'		(Total 6 marks)
	omplete the fompounds.	following statements by ins	serting the names or form	nulae of the elements or
(a)) When carl	bon is heated with copper of	oxide, the products are	
			and	
				(1)
(b)	,	ous elements which, on igremperature are		•
			and	
			and	(1)
(c)		gas can be prepared in the		(1)
(c)	e) Hydrogen	gas can be prepared in the	e laboratory by the react	(1) ion between dilute
(c)	e) Hydrogen		e laboratory by the react	(1) ion between dilute
	e) Hydrogen	gas can be prepared in the	e laboratory by the react	(1) ion between dilute (1)
	e) Hydrogen	gas can be prepared in the	e laboratory by the react	(1) ion between dilute (1)
	e) Hydrogen	gas can be prepared in the	aqueous sodium hydrox	(1) ion between dilute (1) (1)
	e) Hydrogen	gas can be prepared in the	aqueous sodium hydrox	(1) ion between dilute (1) (1)
(d)	Hydrogen Need to be a sulphine with the	gas can be prepared in the	aqueous sodium hydrox	(1) ion between dilute (1) tide, the products are (1)
(d)	Hydrogen Need to be a sulphine with the	gas can be prepared in the	aqueous sodium hydrox	(1) tion between dilute (1) tide, the products are (1)
(d)	Hydrogenl) When sulp	gas can be prepared in the	aqueous sodium hydrox	(1) ion between dilute (1) tide, the products are (1)

Leave blank

6. Complete the table to identify the missing reactant needed to make each gas in the laboratory. Give a test that is commonly used to identify the gas.

Gas	Reactant A	Reactant B	Test
ammonia	ammonium chloride		
sulphur dioxide		dilute hydrochloric acid	
chlorine		concentrated hydrochloric acid	

(Total 6 marks)

Q6



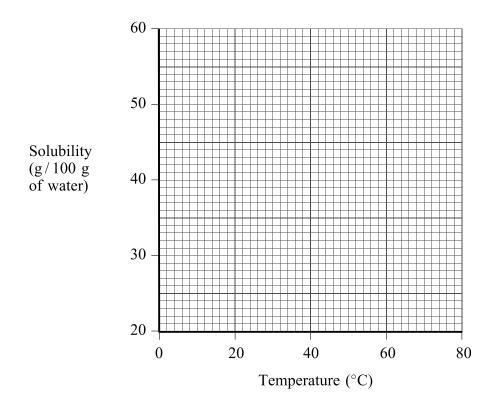
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7.	The	reaction of propane with oxygen is represented by the equation	
		$C_3H_8(g) + 5O_2(g) \rightarrow 3CO_2(g) + 4H_2O(l)$ $\Delta H = -2220 \text{ kJ mol}^{-1}$	
		cm ³ of propane was mixed with excess oxygen and ignited. Assume all asurements are made at room temperature and atmospheric pressure.	
	Cal	culate:	
	(a)	the volume of oxygen that reacted;	
		(1)	
	(b)	the volume of carbon dioxide produced;	
		(1)	
	(c)	the heat energy evolved during the reaction.	
		(3)	Q7
		(Total 5 marks)	

I .	rite a balanced equation for each of the following reactions.	Wri
	The formation of carbon monoxide by the incomplete combustion of carbon.	
	(2)	
	The formation of sodium carbonate and water from aqueous sodium hydroxide and carbon dioxide.	(b)
	(2)	
	The formation of copper, nitrogen and water when copper(II) oxide reacts with ammonia.	(c)
Q8	(2)	
	(Total 6 marks)	

9. The solubility of potassium chloride at various temperatures is given in the table.

Temperature (°C)	0	20	40	60	80
Solubility (g/100 g of water)	28.4	34.2	40.0	45.8	51.3

(a) Use the grid to plot the solubility curve for potassium chloride.



(2)

(b) What is the solubility of potassium chloride at 30 °C?

(1)

(c) If 45 g of potassium chloride is mixed with 100 g of water at 50 °C, what mass of potassium chloride remains undissolved?

(2)

(d) What is the minimum temperature at which 25g of potassium chloride dissolves in 50 g of water?

(4)

(1)

(2)
 (Total 8 marks)
ve the number of:
electrons in the outer shell of selenium (atomic number = 34);
(1)
covalent bonds in one molecule of C_2H_5Br ; (1)
moles of chloride ions that will combine with 1 mole of Ti ³⁺ ions;
(1)
hydrogen atoms in one molecule of the alkane that contains ten carbon atoms;
(1)
moles of atoms in 2 moles of CO(NH ₂) ₂ ;
(1)
moles of electrons needed to liberate 1 mole of bromine molecules during the
electrolysis of molten lead(II) bromide;
(1)
dm ³ occupied by 88 g of carbon dioxide at room temperature and atmospheric
pressure. (1)
(Total 7 marks)

(a) (1)) S	State what is meant by the term hydrocarbon .
		(1)
(ii		Explain, in terms of the bonding, the difference between a saturated and an unsaturated hydrocarbon.
	S	Saturated
	J	Jnsaturated
	•	(2
(b) (i)		Draw the displayed formulae of the saturated and unsaturated hydrocarbon containing three carbon atoms.
		Saturated Unsaturated
		(2)
(ii		On one of the diagrams above, draw a circle around a carbon atom that has its
(ii		
	ii) C	On one of the diagrams above, draw a circle around a carbon atom that has its bonds arranged in a tetrahedral shape.

	Leave blank
(c) Give a structural formula to show the organic product formed when:	
(i) 1 mole of chlorine reacts with 1 mole of ethene;	
(1)	
(ii) 1 mole of chlorine reacts with 1 mole of ethane.	
(1)	Q11
(Total 9 marks)	
	1

	mplete the descriptions of the extraction of the following three metals by inserting the ssing words or formulae.	Leave blank
(a)	Zinc can be obtained from the ore, zinc sulphide, by first heating the ore in air to	
	form zinc; this compound is reduced to zinc metal	
	by heating it in a stream of	
(b)	Chromium can be obtained from its oxide by heating with a metal	
	in the reactivity series. The metal normally used	
	is The equation for the reaction is	
	$Cr_2O_3 + \dots \rightarrow 2Cr + \dots$	
(c)	Aluminium is obtained by the	
	dissolved in molten	
	Aluminium ions electrons at the cathode; the	
	charge on this electrode is	
	The anode is made of	
	anode is which then reacts with this electrode to	
	form gas. (7)	Q12
	(Total 12 marks)	

blank

a) An aqueous solution of a silver salt, A , was added to a solution of a salt, B , a yellow precipitate, C , was formed. When a platinum wire was dipped into sol B and held in a Bunsen flame, a brick-red flame was seen.				
	(3)			
nmonia gas is mixed with air and passed over a catalyst, D , is oxidised to form the gas E and steam. E is cooled and e air to form a brown gas, F . The gas F is reacted with water ne acid G .	is then reacted			
is oxidised to form the gas E and steam. E is cooled and e air to form a brown gas, F. The gas F is reacted with water ne acid G.	is then reacted			
is oxidised to form the gas E and steam. E is cooled and e air to form a brown gas, F. The gas F is reacted with water ne acid G.	is then reacted			
is oxidised to form the gas E and steam. E is cooled and e air to form a brown gas, F. The gas F is reacted with water ne acid G.	is then reacted			
is oxidised to form the gas E and steam. E is cooled and e air to form a brown gas, F. The gas F is reacted with watene acid G.	is then reacted er and more air			
is oxidised to form the gas E and steam. E is cooled and e air to form a brown gas, F. The gas F is reacted with water the acid G.	is then reacted			

Leave	
hlank	

that occurred is represented by the equation $\mathbf{M}(s) + \mathrm{CuSO_4(aq)} \to \mathbf{MSO_4(aq)} + \mathrm{Cu(s)}$ (a) Explain why a hot solution of copper(II) sulphate was used.								
						-		(1)
						(b) E	Explain why an excess of copper(II) sulphate was used.	
-		(1)						
(c) S	Suggest why the copper was dried at 100 °C rather than at a higher temperature.							
-								
-		(1)						
(d) (i) Calculate the number of moles of copper that were formed and hence number of moles of M that were used in the reaction.	the						
		•••••						
		(2)						
(ii) Calculate the relative atomic mass of M .							
		(1)						

	(2)
(f)	Explain, in terms of electron transfer, why the reaction can be considered to involve both oxidation and reduction.
	(2)
	(Total 11 marks)
	TOTAL FOR PAPER: 100 MARKS
	END

	0	2 Helium 4	Neon 20 20 18 Argon	36 Krypton 84	Xeron 131 131 Radon	777		
	_		Fluorine 19 17 17 CI Chlorine 35.5	35 Bromine 80	53 lodine 127 85 At Astatine	2		
	9		Oxygen 16 Sulphur 32					
	2		Nitrogen 14 15 Phosphorus					
	4		Carbon 12 14 14 Silicon Silicon 28	32 Germanium 73	SO 20 119 82 82 634 634 634 634 634 634 634 634 634 634			
	ო		Boron 11 13 Aluminium	31 Ga Gallium 70	hodium 115 115 115 Thallium 126	50		
ш				30 Zinc 65	Cadmium 112 80 Mercury			
THE PERIODIC TABLE				Cu Copper 63.5	Agg Silver 108 79 Agg Agg Agg Agg Agg Agg Agg Agg Agg Ag	<u> </u>		
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	Group	Hydrogen		75	TC Technetium 99 75 Remium 408	8	Key	Symbol Name Relative atomic mass
				٤	Mo Molybdenum T 96 74 W Tungsten	<u> </u>	<u> </u>	
				Ę	Niobium 93 73 Tantalum	0		
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				1 - 1	39 Yttrium 89 57 Lanthanum	Actinium 227		
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