

FIG. 7.1

The mass of lithium in a compound can be found using the equation:

$$\text{mass of lithium (mg)} = \frac{\text{mass of compound (mg)} \times \text{number of ions of Li in formula} \times A_r \text{ of Li}}{M_r \text{ of compound}}$$

where A_r of Li = 6.9 and M_r = relative formula mass of a lithium compound.